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AS Chemistry

C4: Inorganic Chemistry and Periodic Table -2

Assignment Questions

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Year 2024

Inorganic Chemistry and Periodic Table -2

Q1: Explain the trends in boiling temperature down group 7.

Q2: Bromine reacts with hot concentrated sodium hydroxide solution to form sodium bromate (V) , sodium bromide and water. The reaction is similar to the one between chlorine and hot concentrated sodium hydroxide solution.

a) Write the equation for the reaction above.

b) State the type of reaction and explain your answers in terms of oxidation numbers.

Q3: Silver chloride and silver bromide both are solids at room temperature.

a) State the colour of both the compounds.

b) Explain how aqueous ammonia can be used to distinguish between the two solids.

Q4: What are the observations and products formed when sodium iodide is reacted with concentrated sulfuric acid?

Q5: Bromine and potassium iodide reacts in a solution to produce potassium bromide and iodine.

a) Write the equation for this reaction.

b) Write an ionic equation for this reaction.

c) What would you see if some cyclohexane was added to the mixture , then shaken?