



KEC

K Education Centre



AS Chemistry

C2: Chemical Bonding and Structure Assignment -1

Assignment Questions

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Chemical Bonding and Structure Assignment -1

Q1 : Describe metallic bonding.

Q2: Describe ionic bonding.

Q3: Why magnesium ion Mg^{2+} will make a strong ionic bond compare to Li^+ ?

Q4: What are the factors affecting the strength of ionic bonding?

Q5: Which of the following ionic compounds will required more energy to separate the ions. Give reasons.

a) LiF

b) KF

Q6: What are isoelectronic ions? Give examples.

Q7: What are general trends for ionic radii

a) Across a period

b) Down the group

Q8: Why Lattice energy (measure of strength of ionic bond) of lithium oxide

Li_2O is more than Potassium Oxide K_2O

Q9 : Why ionic compounds have high melting points but are brittle.

Q10: What is covalent bonding?

Q11: What are sigma and pie bonds? Give examples.

Q12: Why alkenes are more reactive than alkanes?

Q13: Draw dot and cross diagram to show the bonding in

a) Ethyne C_2H_4 b) NH_3

Q14: Why Cl - Cl is more stronger than I - I ?